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A gas is a state of matter with no defined shape or volume. Gases have their own unique behavior depending on a variety of variables, such as temperature, pressure, and volume. While each gas is different, all gases act in a similar matter. This study guide highlights the concepts and laws dealing with the chemistry of gases.

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- To describe the particle nature of both real and ideal gases.
- To describe the properties of gases that can be used to explain their characteristics: volume, number of particles, temperature, and pressure.

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Effusion is the gas movement through a tiny hole one particle at a time. Small, light gases have greater speeds than large, heavy gases at a given temperature, so they effuse faster. Graham's Law of Effusion relates effusion to mass. Chapter 13 "Gases" What are the properties of gases?

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Equal volumes of gases at the same temperature and pressure contain equal numbers of particles. b. One mole of any gas will occupy a certain volume at STP. c. STP stands for standard temperature and pressure. d. The molar volume of a gas is the volume that one mole occupies at STP. temperature volume pressure 5. 6. 4. CHAPTER 13 STUDY GUIDE

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